1.13B - Standard Atomic Notation and Sub-atomic Particles for the First 20 Elements

Instructions:

1. Print the **standard atomic notation** for the element at **the top of its box** as shown for **hydrogen** below.

Important: To find the mass number, round the **atomic mass** of the element from your Periodic Table to the **nearest whole number**.

2. The number of **protons** (p^{+1}) and **electrons** (e^{-1}) for each element are both equal to the **atomic number**. Write them in the space provided.
3. Calculate the number of neutrons (n^0) for each element. The number of **neutrons** is equal to the **difference between the mass number and the atomic number.** Write the number of neutrons from your calculation in the space provided for each element.

$^{1}_{1}H$	$^{4}_{2}He$
# p^{+1} = 1	# p^{+1} = 2
# e^{-1} = 1	# e^{-1} = 2
# n^0 = 0	# n^0 = 2

$^{7}_{3}Li$	$^{9}_{4}Be$	$^{11}_{5}B$	$^{12}_{6}C$	$^{14}_{7}N$	$^{16}_{8}O$	$^{19}_{9}F$	$^{20}_{10}Ne$
p 3	4	5	6	7	8	9	10
e 3	4	5	6	7	8	9	10
N 4	5	6	6	7	8	10	10
$^{23}_{11}Na$	$^{24}_{12}Mg$	$^{27}_{13}Al$	$^{28}_{14}Si$	$^{31}_{15}P$	$^{32}_{16}S$	$^{35}_{17}Cl$	$^{40}_{18}Ar$
p 11	12	13	14	15	16	17	18
e 11	12	13	14	15	16	17	18
N 12	12	14	14	16	16	18	22
$^{39}_{19}K$	$^{40}_{20}Ca$						
p 19	20						
e 19	20						
N 20	20						